

VIDEO PEMBELAJARAN

MATERI: CHEMICAL KINETIC

PART 4

Dapat diakses di: <http://youtu.be/yf2GArmVCxM> atau
<http://jeikhsan.wordpress.com/kuliah/>

The screenshot shows a YouTube video player for 'Jaslin Ikhsan: Chemical Kinetic Part 4'. The video is paused at 04:18 / 11:35. The main content is a slide titled 'Outline: Kinetics' with the following table:

	First order	Second order	Second order (Bimolecular)
Rate Laws	$rate = -k[A]$	$rate = -k[A]^2$	$rate = -k[A][B]$
Integrated Rate Laws	$\ln \frac{[A]_t}{[A]_0} = -kt$	$\frac{1}{[A]_t} = kt + \frac{1}{[A]_0}$	complicated (discussed later)
Half-life	$t_{1/2} = \frac{0.693}{k}$	$t_{1/2} = \frac{1}{k[A]_0}$	Complicated (discussed later)

Below the video player, it says 'Uploaded by jaslinkhsan on May 27, 2011' and 'Other order (nth order) Reaction: Rate Laws, integrated rate laws'. The video has 24 views and 0 likes/dislikes. The right sidebar shows a list of related videos, including 'Jaslin Ikhsan: Chemical Kinetic Part 1', 'CHE343 - Chemical Kinetics and Reactor Design', and 'Jaslin Ikhsan: Collision of Gas Molecules Part ...'.

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Second-Order Processes

A graph of $1/[NO_2]$ vs. t gives this plot.

$$\frac{1}{[A]_t} = kt + \frac{1}{[A]_0}$$

- This is a straight line. Therefore, the process is second-order in $[NO_2]$.

Time (s)	$[NO_2], M$	$1/[NO_2]$
0.0	0.01000	100
50.0	0.00787	127
100.0	0.00649	154
200.0	0.00481	208
300.0	0.00380	263

02:41 / 11:35 3672

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Other order (nth order) Reaction: Rate Laws, integrated rate laws 0 likes, 0 dislikes Show more

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Chemical Kinetics - Concept

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