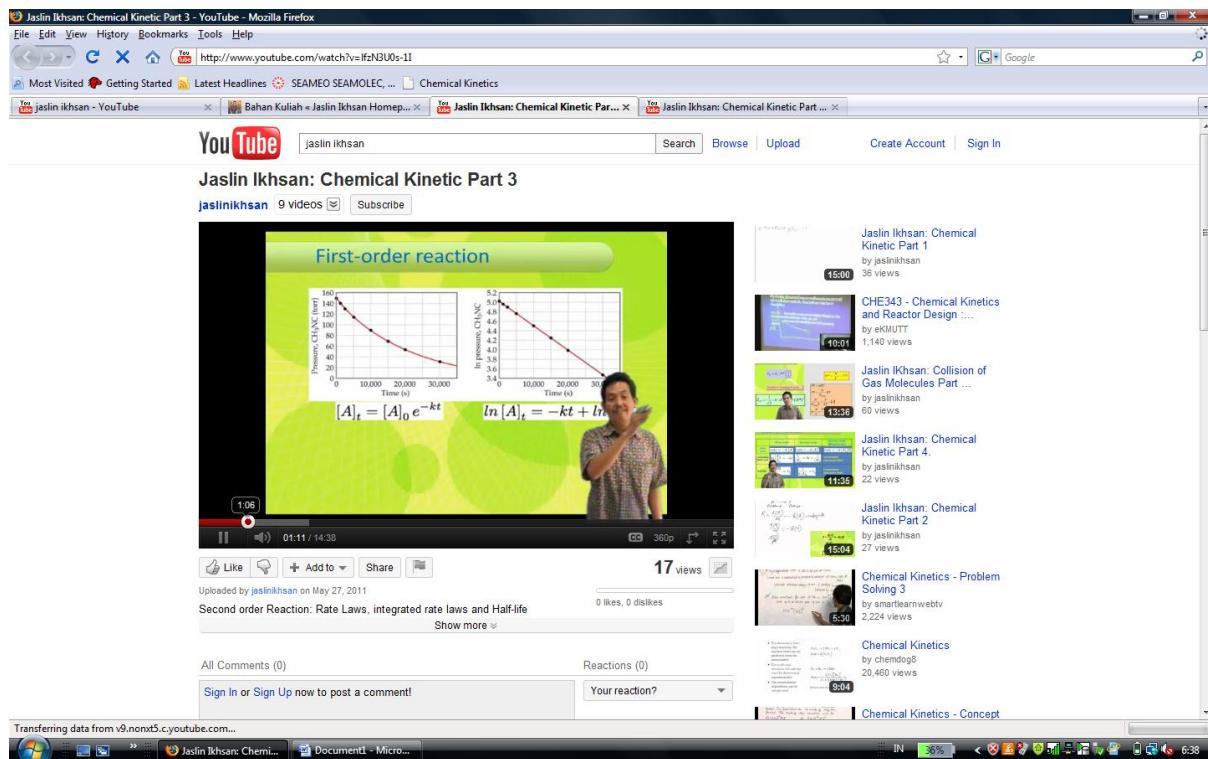


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PART 3

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1st order rxn.
 $\ln\left(\frac{[A]}{[A_0]}\right) = -kt \rightarrow \ln\left(\frac{[A]}{[A_0]}\right) = -kt$
 $\Rightarrow [A] = [A_0]e^{-kt}$
 $\ln\left(\frac{[A]}{[A_0]}\right) = -kt$
 $\ln\left(\frac{0.5}{1}\right) = -kt$
 $0.693 = kt$

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Second order Reaction: Rate Laws, integrated rate laws and Half-life

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